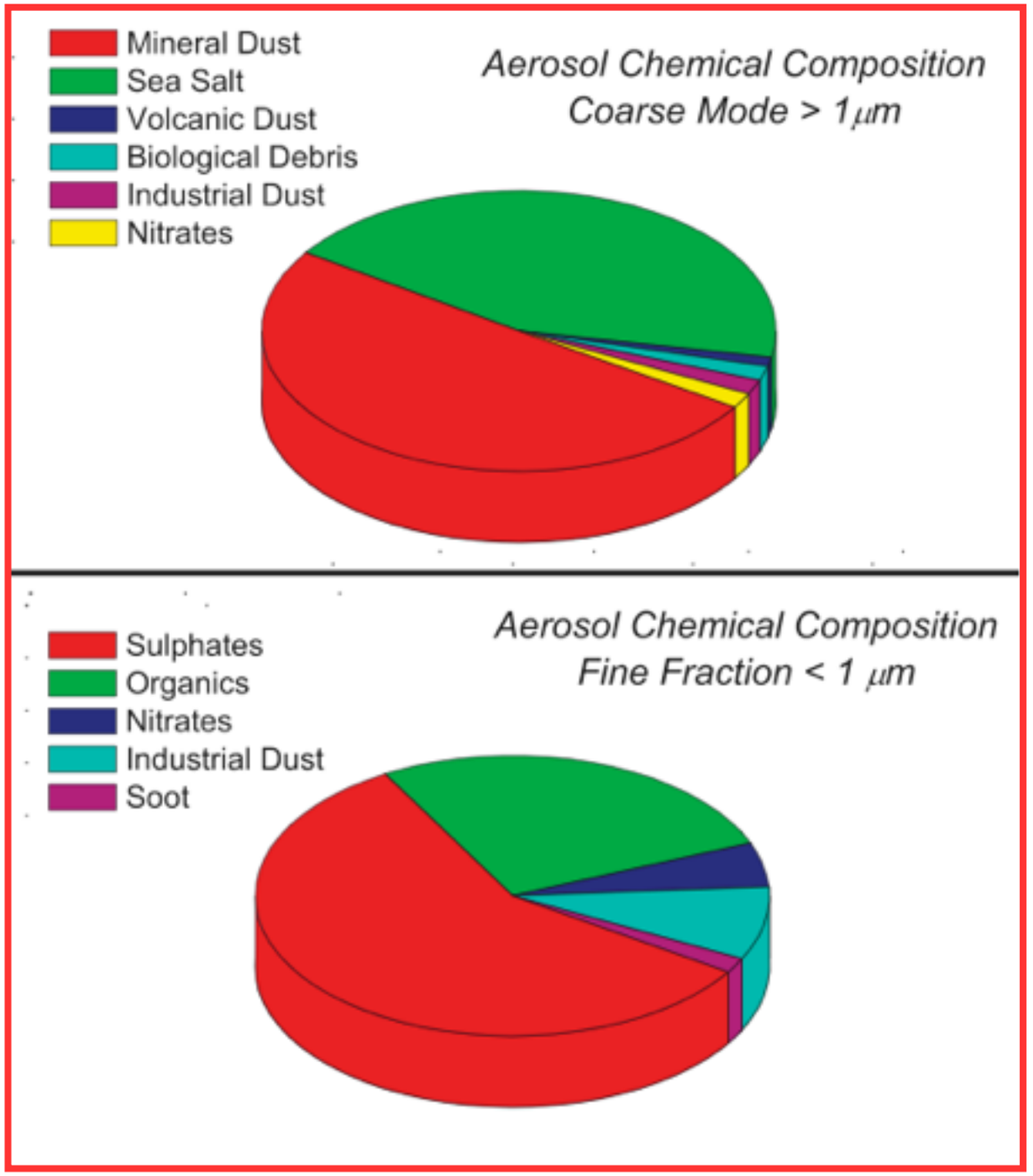
Heterogeneous chemistry - how to get from the literature to UKCA model treatment

Paul Griffiths

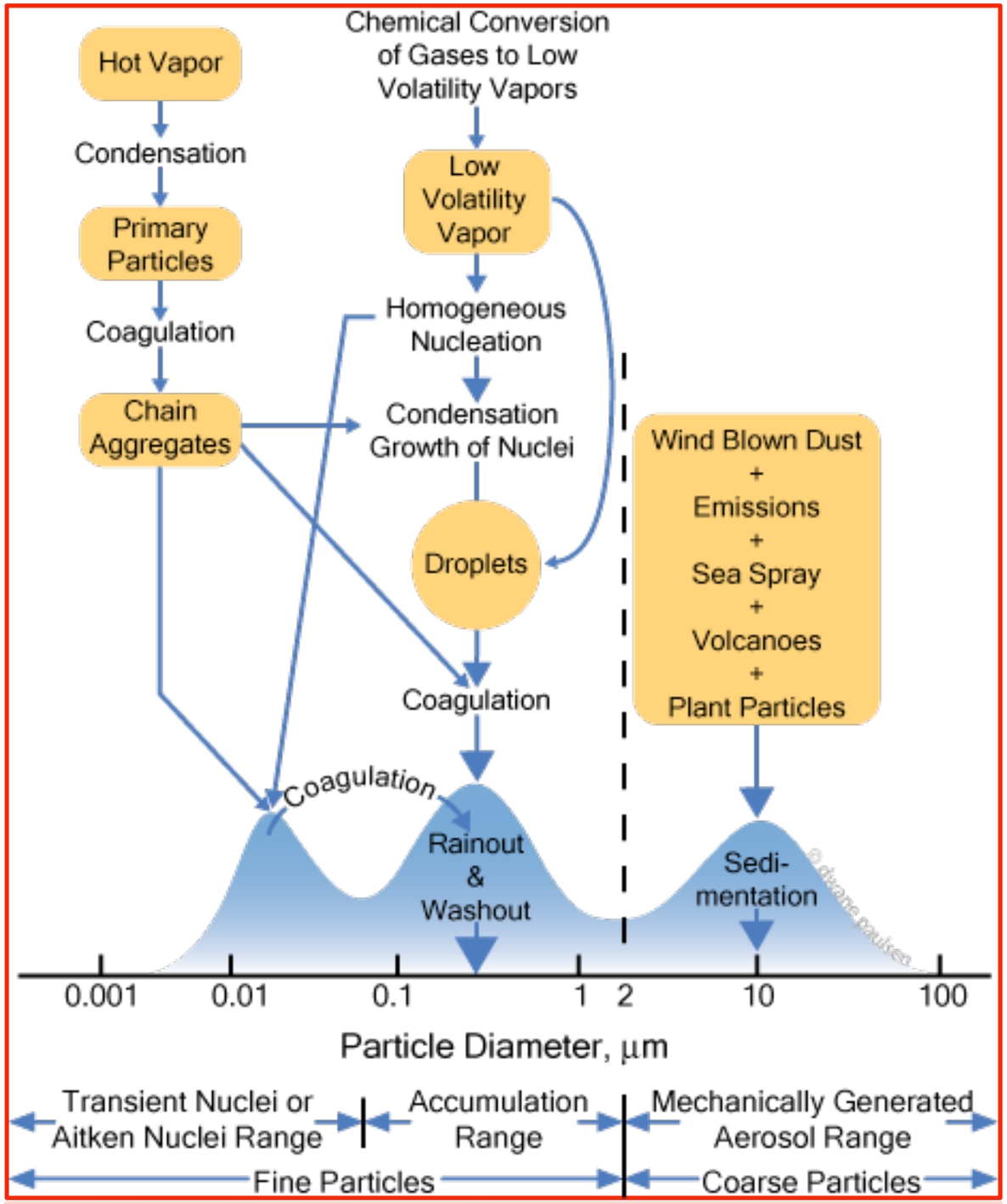
Chemical composition of tropospheric aerosol

- The particulate matter in the atmosphere is very diverse.
- Wide variety of primary sources - dust and sea salt are the strongest by mass.
- Chemical transformation of pollutants occurs within the atmosphere, and can lead to aerosol formation.
- $NO_x \rightarrow HNO_3 \rightarrow$ aerosol nitrate
- $SO_2 \rightarrow H_2SO_4 \rightarrow$ aerosol sulfate
- VOC → organics (carbon-rich)

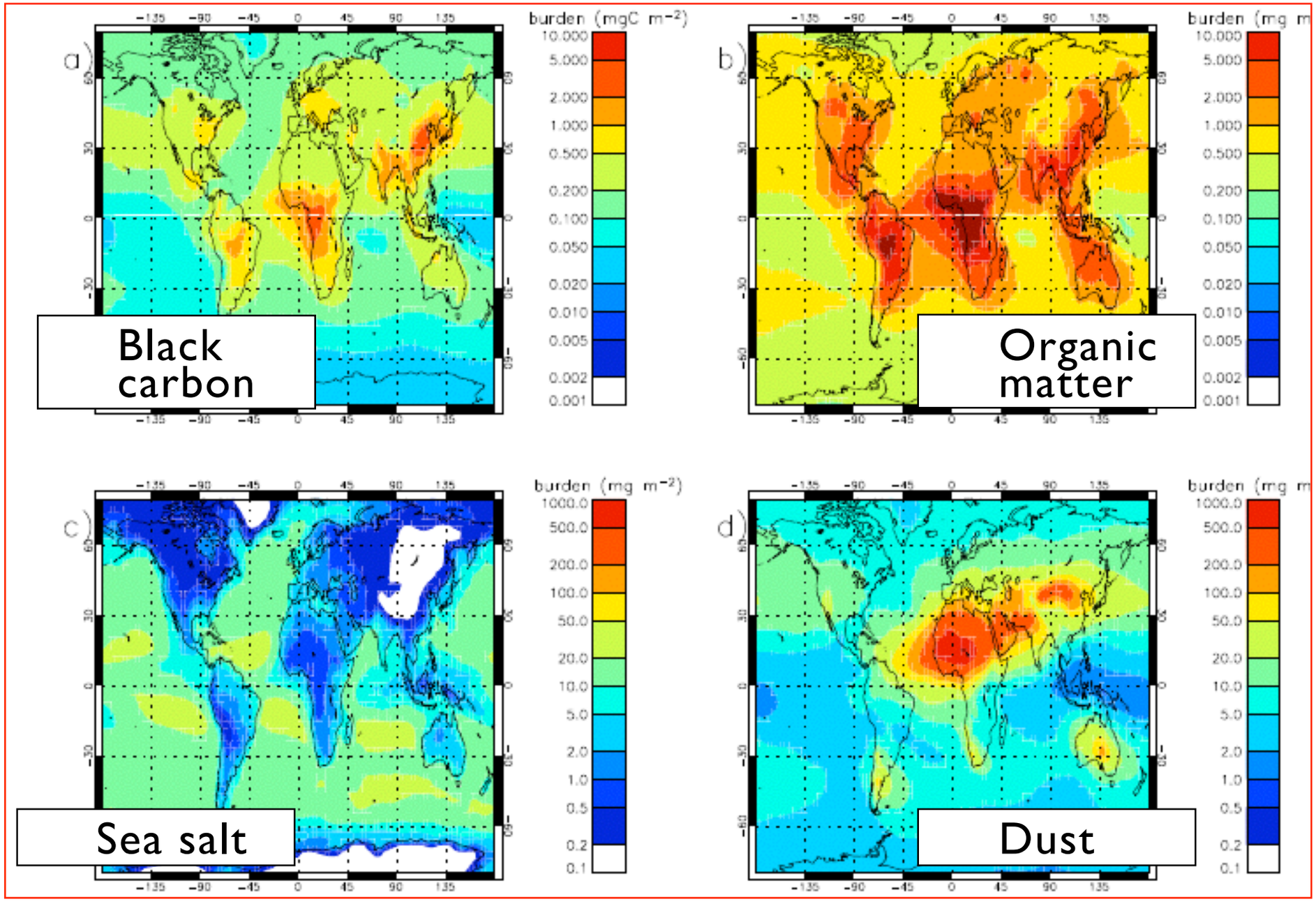


Size range spanned by atmospheric aerosol

- Sizes range from few nm up to hundreds of microns.
- The lifetime of small and large particles is short, but between 100nm and 2 μm, the lifetime can be up to a couple of weeks.
- Dust can be transported across oceans, volcanic aerosol can survive in the stratosphere for years.
- The integrated impact can be large.

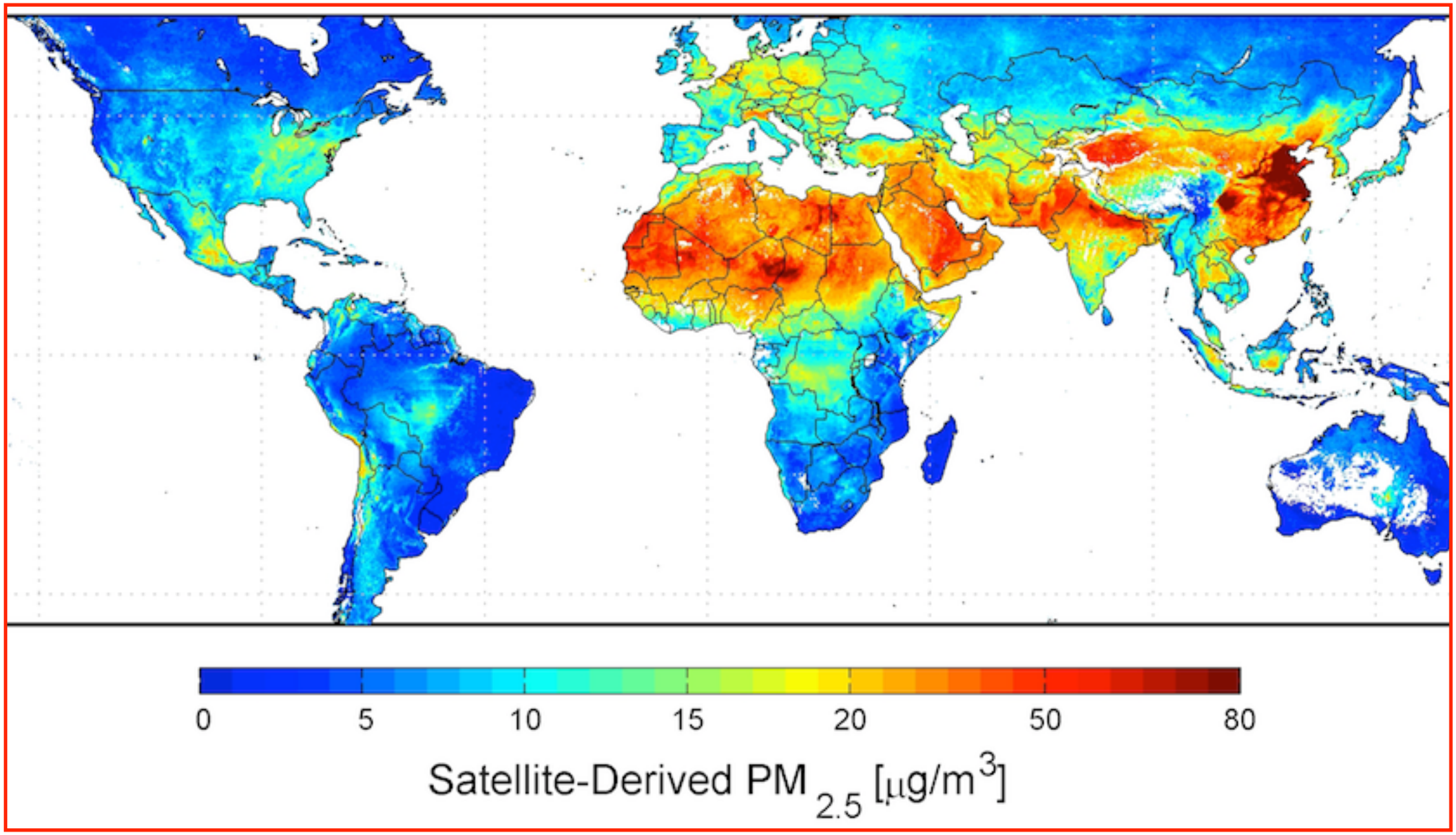


Current work - aerosol chemistry on a global scale



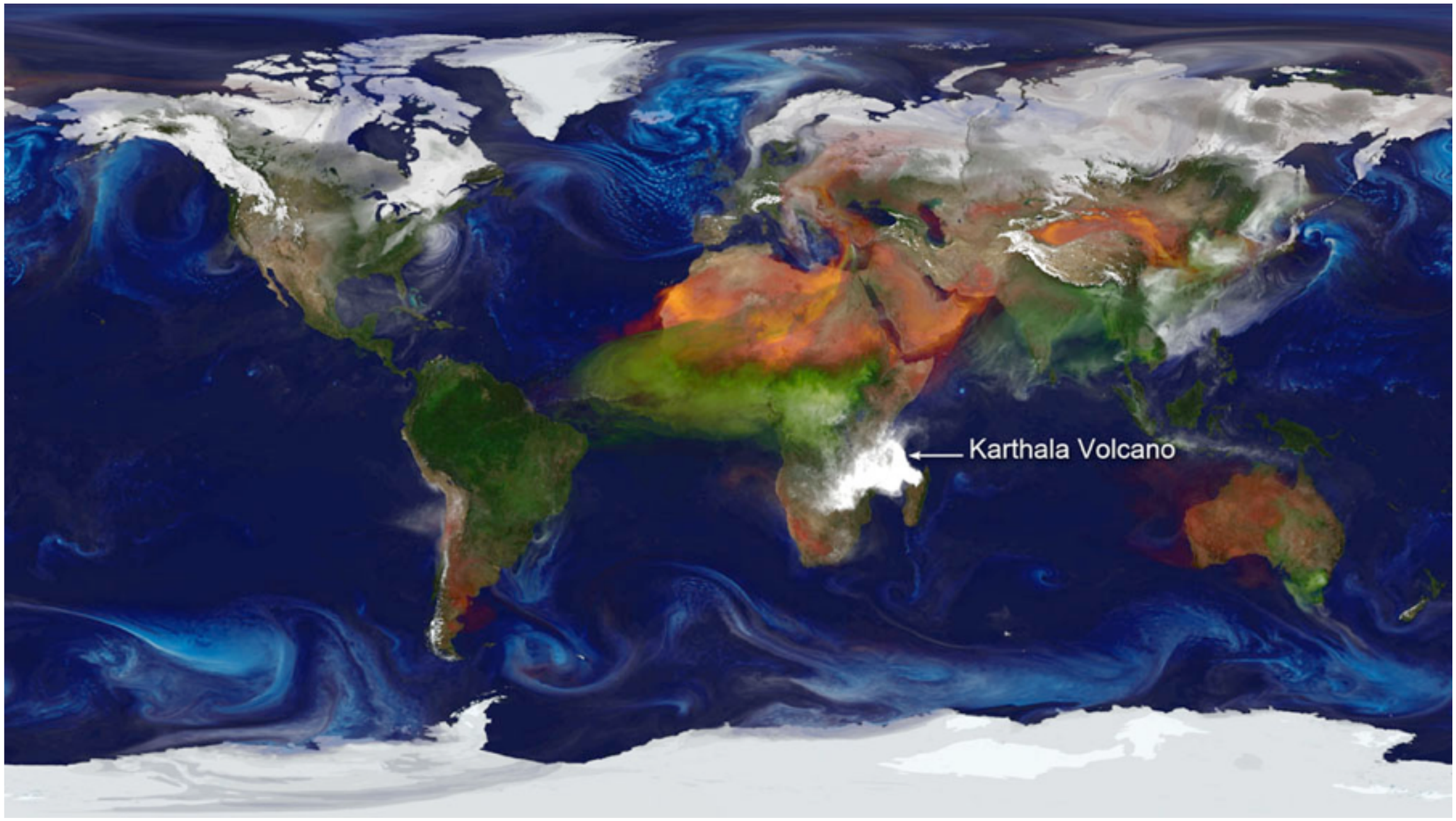
This aerosol model runs within the UK Met Office 'Unified' Climate Model.

Global near-surface distribution of aerosol



Total (mass) particulate matter is shown (NB no measurements over oceans)

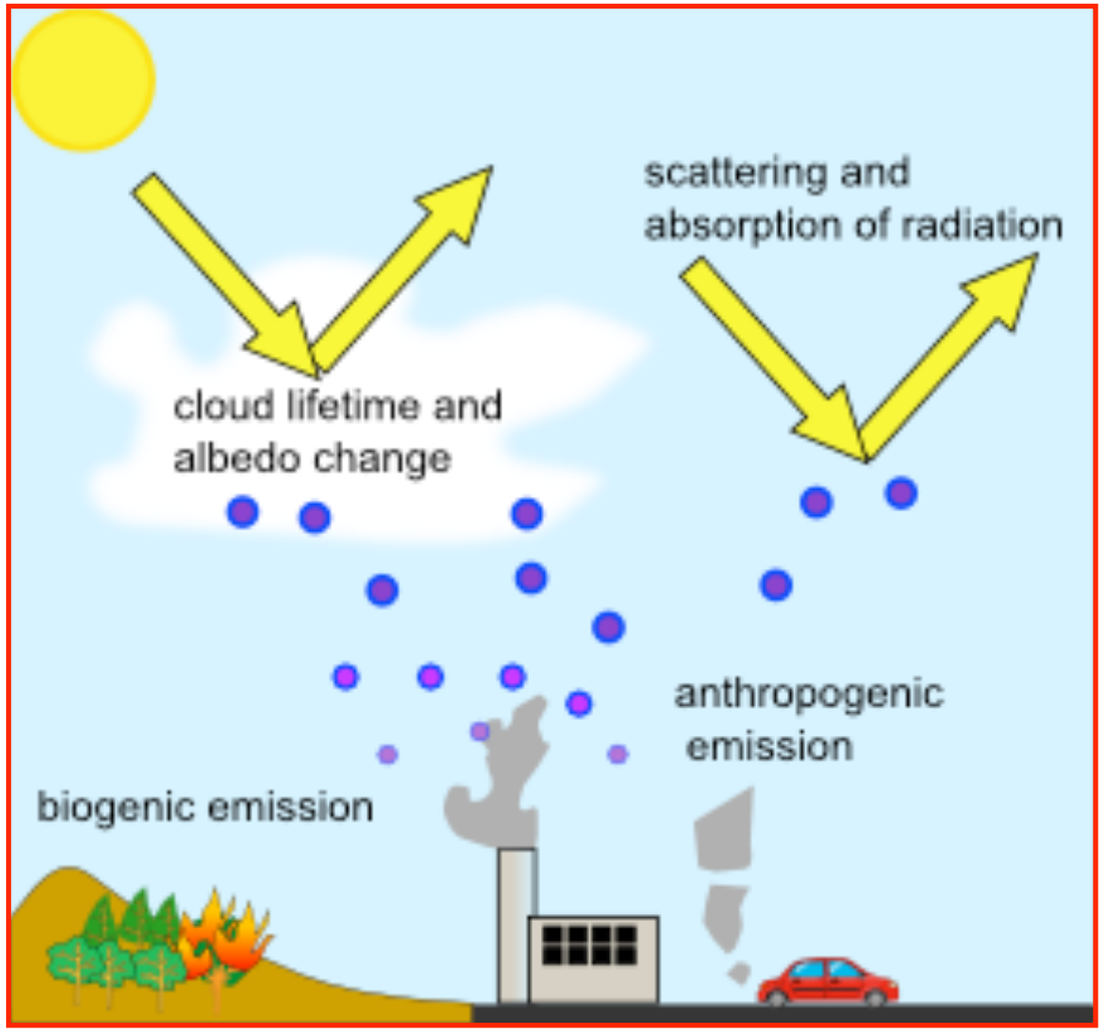
Global distribution of aerosol



White: sulfate blue: sea salt red: dust green: soot

Aerosols and the chemistry of the troposphere

- Anthropogenic pollutants are degraded by the action of sunlight in the atmosphere to form aerosol.
- Absorption and scattering by the aerosol then change the amount of incident sunlight - local climate different in the neighbourhood of the emissions.
- Also changes the rate of degradation - a feedback loop is established.



Coupling between aerosol and oxidation

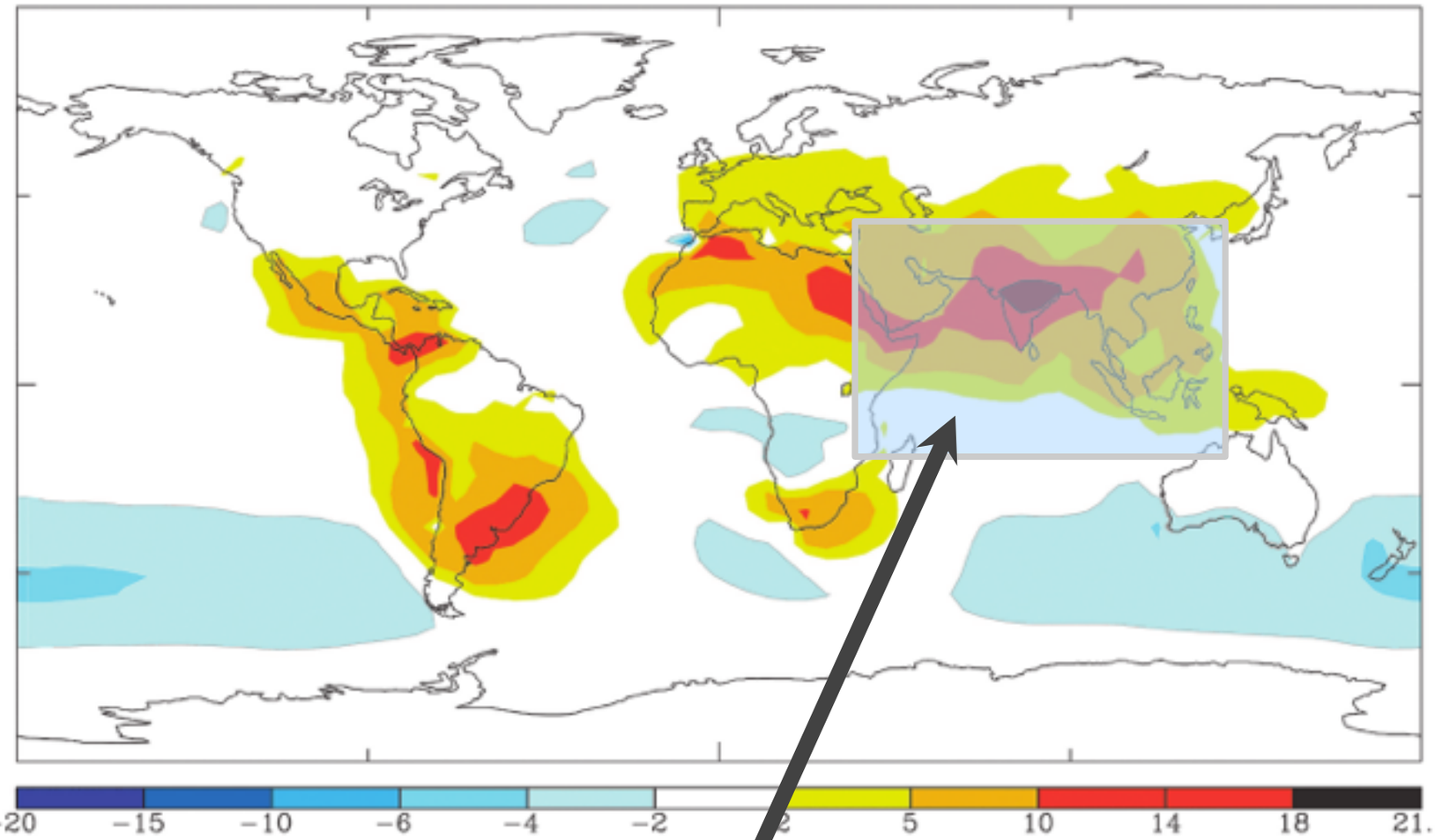


Fig. 1. Percentage difference in surface sulfate mixing ratio between future control simulation (2030C) and sensitivity simulation with future SO_2 emissions but present-day O_3 precursor emissions (2030SO_4) [i.e., 100 × (2030C − 2030SO_4)/2030C].

Regional variations in impact: higher future NOx emissions lead to large increase in [OH], more OH+SO_2, more sulfate. More aerosol means a decrease in solar radiation, moderating climate change.

Aerosols and their effect on ozone

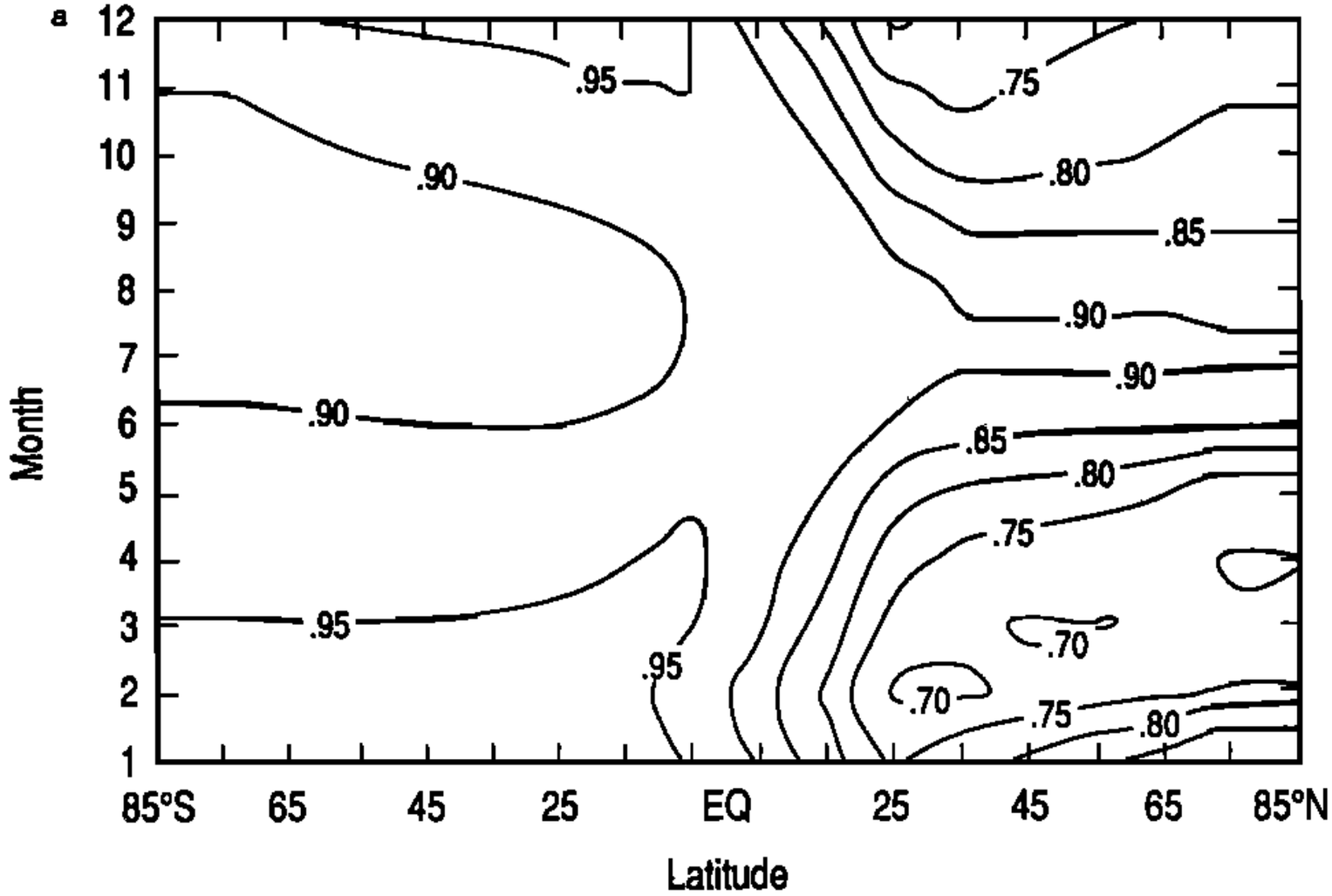


Fig. 9 from Dentener and Crutzen, 1993. The zonal mean monthly average ratio of O3 concentrations with aerosol reactions to those without for all months at 1000 hPa

Particulate matter in the atmosphere

- Serve as the germ or nucleus for cloud formation.
- Scatter incoming solar radiation (reduce heating at the surface).
- Absorb outgoing radiation (increase strength of greenhouse effect).
- Remove (and transform) trace gases and pollutants.
- Health effects

Chemical effects of particles in the atmosphere

- Partitioning to the aqueous phase - solubility controlled
- Adsorption onto solid surfaces
- Reactive uptake of trace gases
- Release of trace gases following reactive uptake

- **Aerosols may enhance concentration and so open up new pathways of reactivity.**

Prognostic equation for reaction of aerosol with gas X

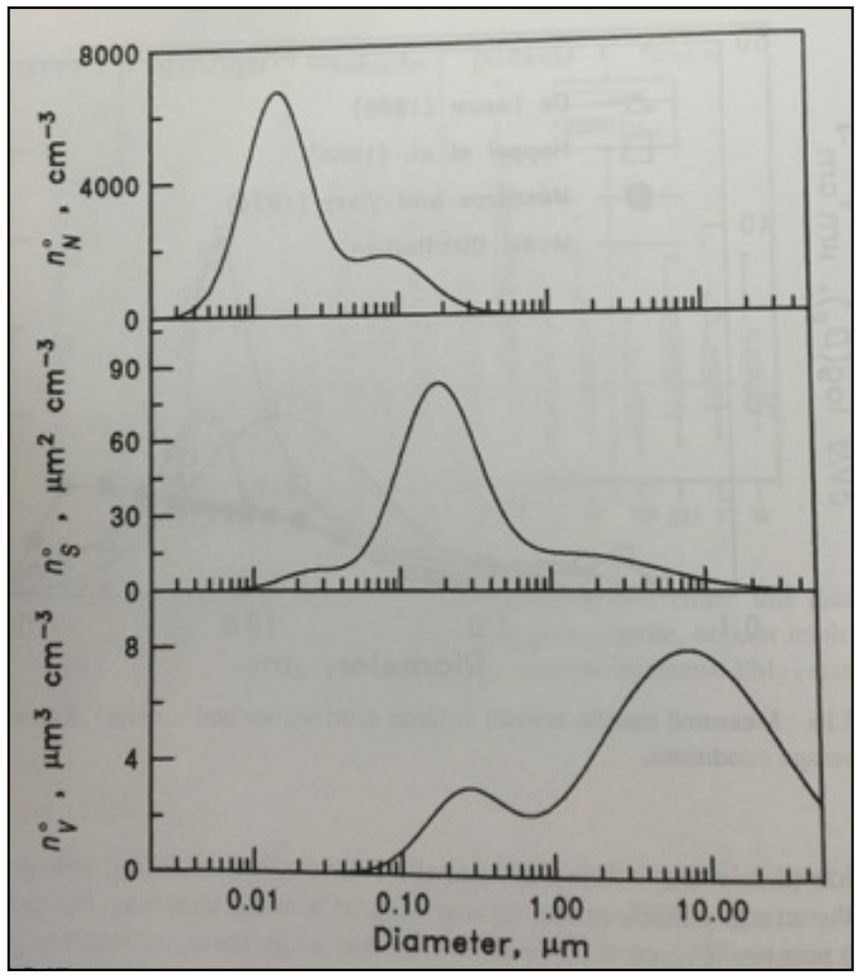
$$\frac{d[X]}{dt} = -k_{het}[X] \qquad c = \sqrt{\frac{8kT}{\pi m}}$$

$$k_{het} = \gamma S_a c/4$$

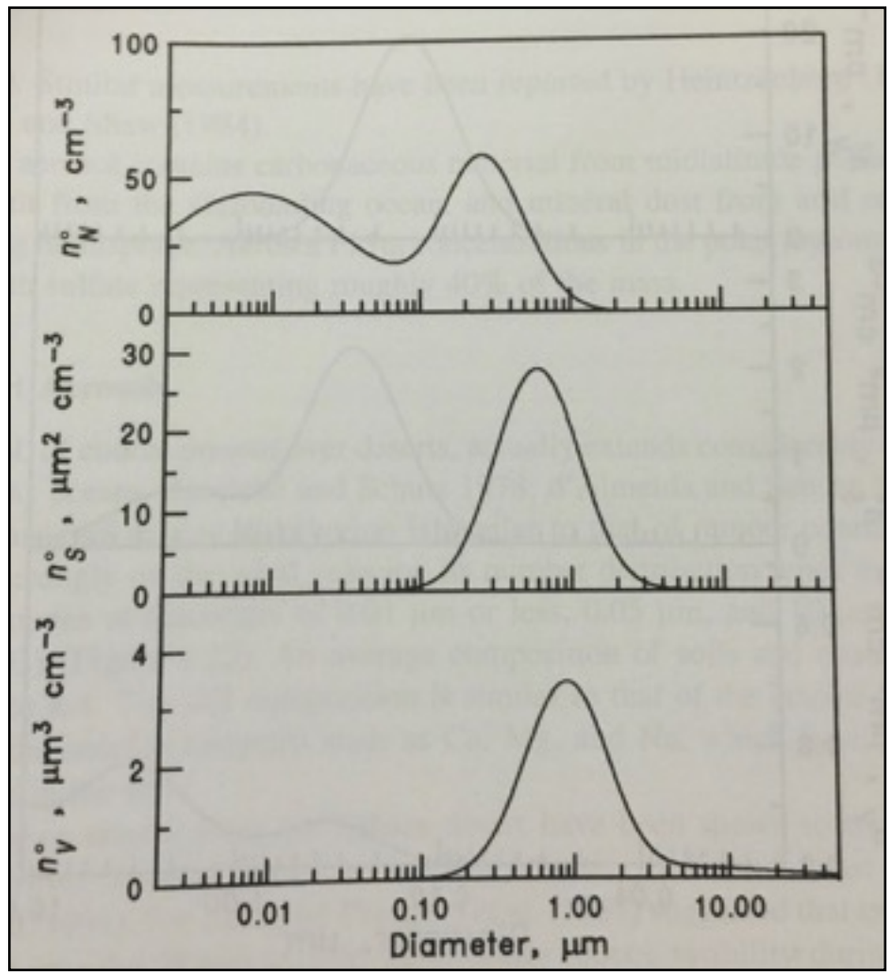
- Parameterise reaction/uptake via uptake coefficient, γ, ($0 < \gamma \leq 1$)
- Represents the probability that the gas phase species X is lost per collision with the aerosol surface.
- Make sure units of surface area density ($m^2\ m^{-3}$) and c (mean molecular speed of gas phase X, ms^{-1}) match.
- E.g. typical units of SA density - $\mu m^2\ m^{-3}$

Representative aerosol number, SA and mass distributions

Rural continental

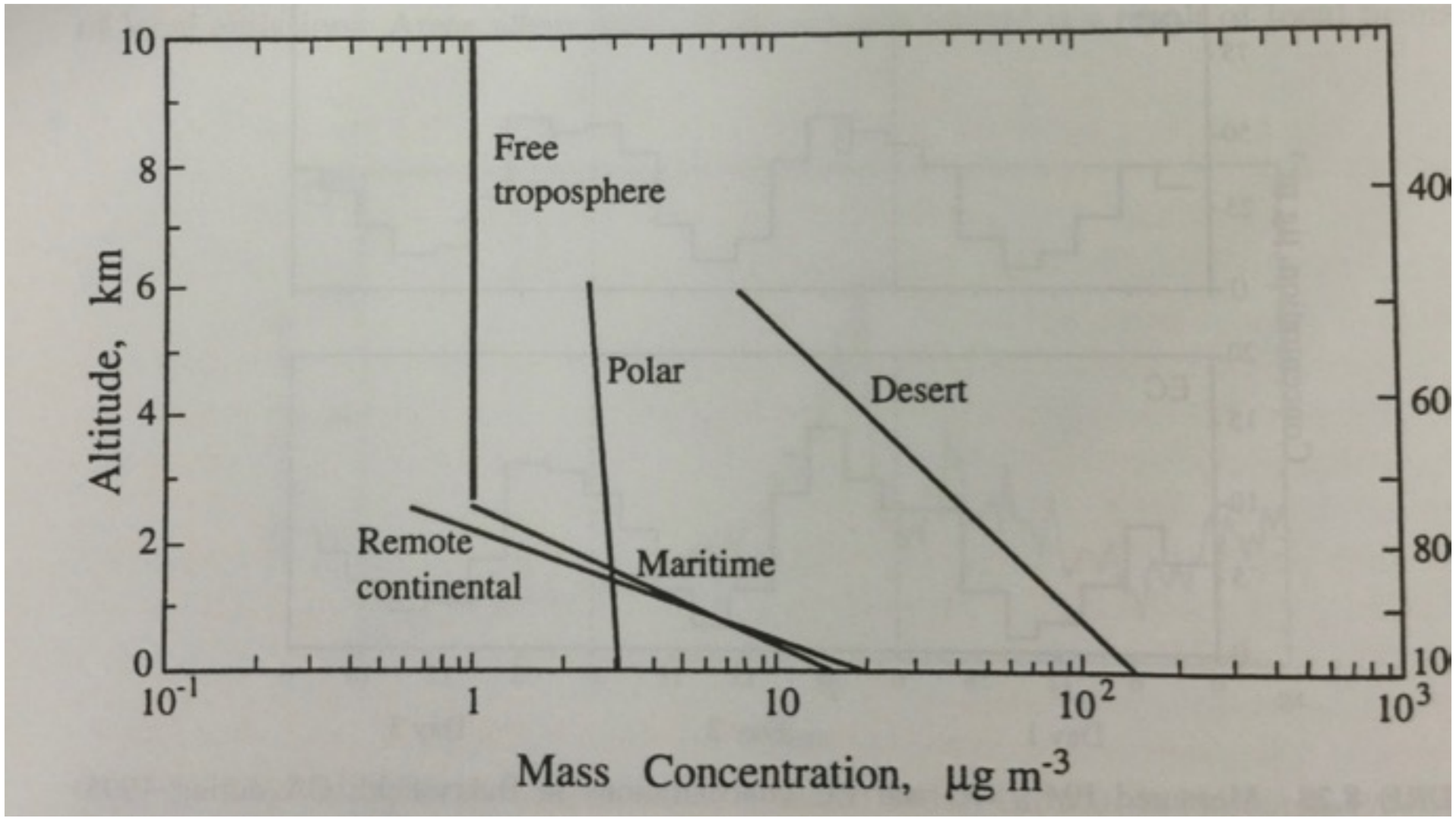


Free troposphere



Maximum of aerosol surface area density is in the accumulation mode

Tropospheric profile of aerosol mass



Figures taken from Seinfeld and Pandis, p. 375

Kinetics of removal of trace gases

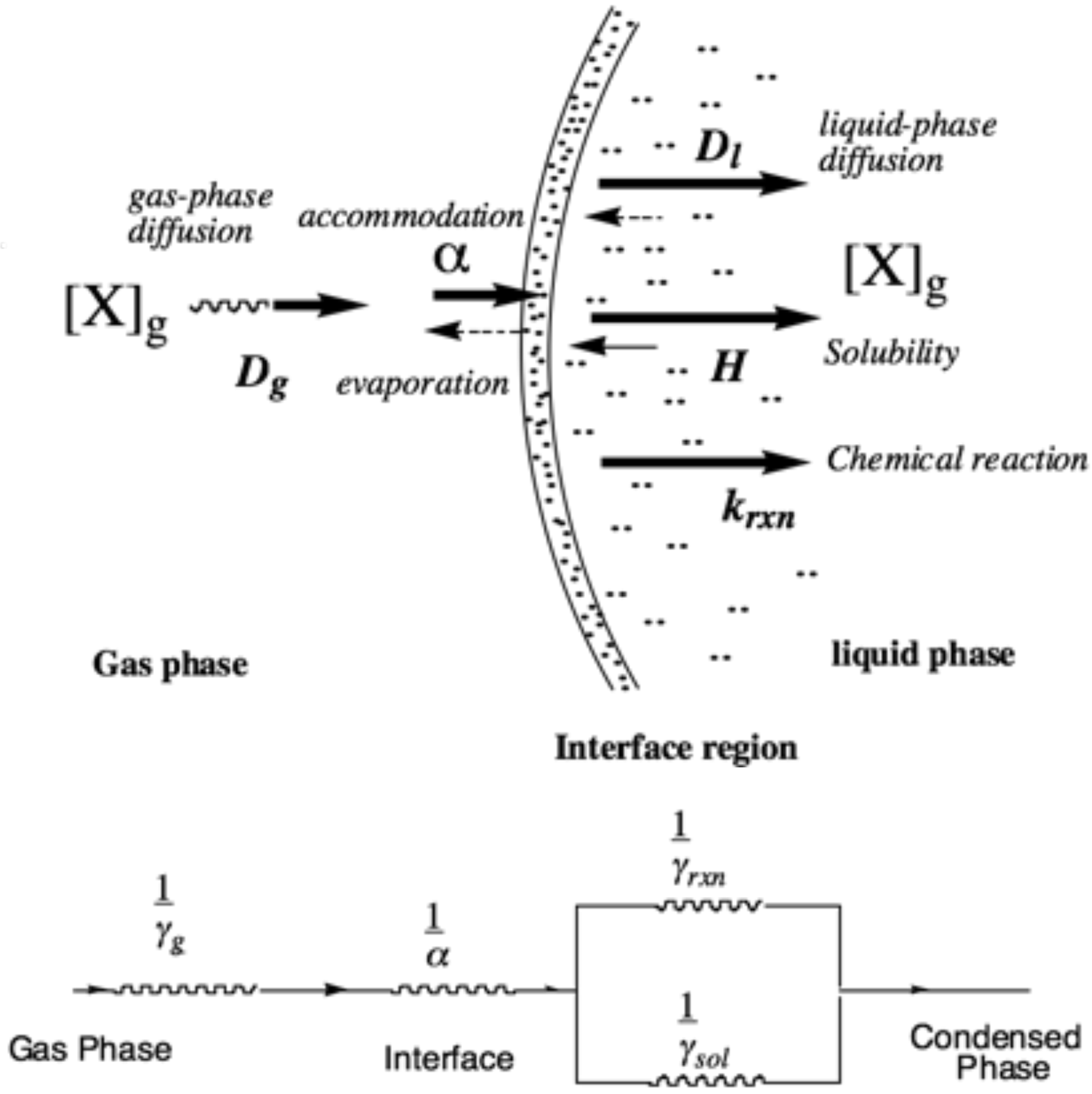
$$\frac{d[X]}{dt} = -k_{het}[X]$$

$$k_{het} = \gamma S_a c/4$$

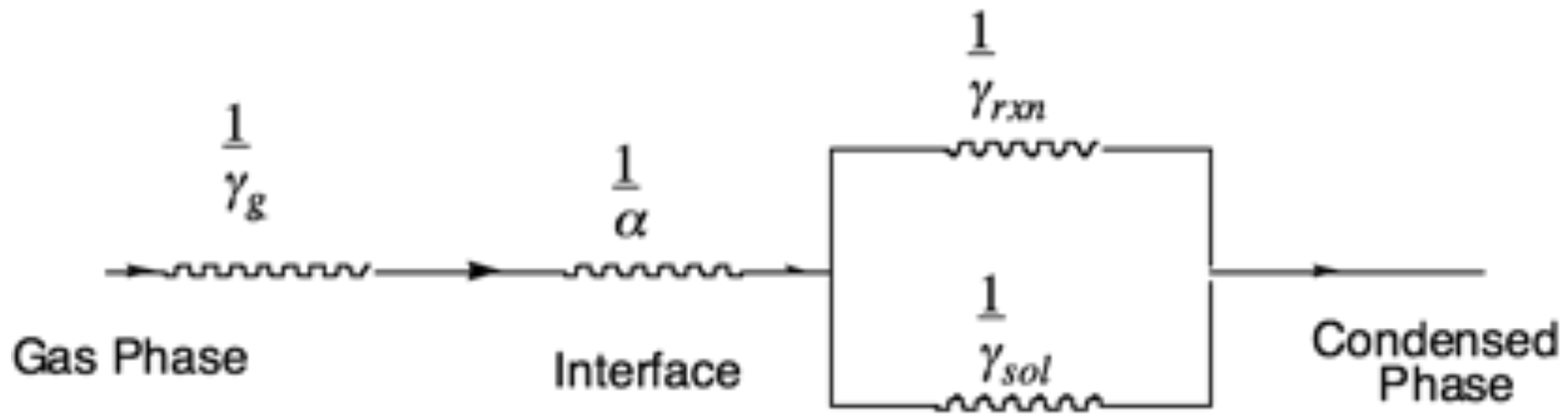
$$c = \sqrt{\frac{8kT}{\pi m}}$$

- The uptake coefficient, γ, contains the physical chemistry that you need to think about.
- This is generally what the experimentalists report.
- Tempting just to take the number from the paper and put it into the model, but there's usually scope to add a bit of value.
- The resistor framework provides the best way currently to do this, although there are other modelling frameworks available - e.g. Shiraiwa (PRA framework).

Trace gas uptake by liquids – a series of coupled steps



Trace gas uptake as a series of coupled steps



$$\frac{1}{\gamma} = \frac{1}{\gamma_g} + \frac{1}{\alpha} + \frac{1}{\gamma_{sol} + \gamma_{rxn}}$$

- The value of gamma can be set by any one of these processes.
- smallest $\gamma \Rightarrow 1/\gamma$ dominates sum - pays to identify the slowest step!
- If no reaction($\gamma_{rxn} = 0$) , then expect the aerosol to become saturated in the trace gas of interest. Over time, the net uptake will reduce to zero.

Trace gas uptake as a series of coupled steps

$$\frac{1}{\gamma} = \frac{1}{\gamma_g} + \frac{1}{\alpha} + \frac{1}{\gamma_{sol} + \gamma_{rxn}}$$

- γ_g represents the modification to the collision frequency (and uptake rate) if the particle is large or diffusion is slow to the particle surface. Depends of particle radius,r, and D_g, gas phase diffusion coefficient

$$\frac{1}{\gamma_g} = \frac{cr}{4D_g}$$

- α is the probability that having collided with the particle surface, the molecule X enters or sticks to the aerosol (NB one way). This can be the smallest probability for solid particles for which this probability is low and so this may control gamma.

- Conversely, when reaction is very fast (γ_{rxn} is large) the accommodation step may again be the slowest step and control γ (so-called 'accommodation-controlled uptake')

Trace gas uptake as a series of coupled steps

$$\frac{1}{\gamma} = \frac{1}{\gamma_g} + \frac{1}{\alpha} + \frac{1}{\gamma_{sol} + \gamma_{rxn}}$$

- γ_{sol} represents the uptake due to the equilibration of the condensed and gas phases - establishing a Henry's law equilibrium at which uptake and evaporation are equal - net uptake at this point is zero.

- Key points: that solubility is composition dependent - e.g. solubility of SO_2 increases in acidic aerosol; aerosol has probably equilibrated in a UKCA timestep.

$$\frac{1}{\gamma_{sol}} = \frac{\sqrt{\pi}c}{4HRT}\left(\frac{t}{D_l}\right)^{1/2}$$

- Fraction of gas X in aqueous phase can be calculated as

$$\frac{[X_{aq}]}{[X]} = 10^{-6}HRTL$$

- *L* is liquid water content in g cm^{-3}. Can be significant for soluble gases in clouds!

Trace gas uptake as a series of coupled steps

$$\frac{1}{\gamma} = \frac{1}{\gamma_g} + \frac{1}{\alpha} + \frac{1}{\gamma_{sol} + \gamma_{rxn}}$$

- γ_{rxn} represents the uptake due to reaction within the aerosol volume.
- Key points: is composition dependent - N_2O_5 onto sulphate (fast) vs nitrate (slow); HO_2 onto organic aerosol; O_3 uptake by SOA. Most general form (rarely used)

$$\frac{1}{\gamma_{rxn}} = \frac{c}{4HRT\sqrt{D_l k_{rxn}}}\left(\coth q - \frac{1}{q}\right)$$

- krxn is the first-order (s^{-1}) rate constant for loss of X in solution; q is the ratio of particle radius to reacto-diffusive length, l.

$$q = \frac{r}{l} \qquad\qquad l = \sqrt{\frac{D_l}{k_{rxn}}}$$

- In general, the above equation can be simplified.

Uptake onto solid surfaces

$$\frac{1}{\gamma} = \frac{1}{\alpha} + \frac{1}{\gamma_s}$$

- Here reaction is between a surface site or surface bound molecule
- Two limiting cases - Langmuir-Hinshelwood (most common by far) or Eley-Rideal
- For Langmuir-Hinshelwood, the reaction is between gas-phase X and a surface-bound molecule Y (here Y_s). Need info on concentration of Y and on partitioning at equilibrium between X and surface-bound X, given by K and N.

$$\frac{1}{\gamma} = \frac{1}{\alpha} + \frac{c(1 + K[X])}{4k_{rxn}[Y_s]KN}$$

- For more information see the IUPAC website, http://iupac.pole-ether.fr/htdocs/supp_info/NewHetIntroOct2009.pdf

Back of the envelope

- Need a representative surface area density (or range)
- And a gamma (or range of gammas)
- Then calculate k_{het} for the trace gas of interest
- Compare with other relevant rates e.g. reaction with O_3, dry deposition, wet removal, photolysis, …
- Can also put into a box model and do an aerosol on/off experiment
  - E.g. KPP - http://people.cs.vt.edu/~asandu/Software/Kpp/
  - MECCA - http://www.rolf-sander.net/messy/mecca/mecca_box.html
  - DSMACC - https://github.com/barronh/DSMACC

Putting heterogeneous chemistry into UKCA

- Follow the tutorial for putting in a new chemical reaction.
- http://www.ukca.ac.uk/wiki/index.php/UKCA_Chemistry_and_Aerosol_Tutorial_6#Heterogeneous_Reactions

Heterogeneous reactions are those that occur on aerosol surfaces. There is no functional form defined for these reactions, with special code needed to be added for each case.

Heterogeneous Reaction Definition

The heterogeneous reactions are defined in the **ukca_chem_scheme.F90** routines using the **rath_t** Fortran type specification, usually in one array. To format of this **rath_t** type is

```
rath_t('Reactant 1','Reactant 2','Product 1 ','Product 2 ','Product 3 ',&
'Product 4 ', Fraction of Product 1 produced, Fraction of Product 2 produced, &
Fraction of Product 3 produced, Fraction of Product 4 produced), &
```

i.e. there is no rate information provided. For reactions on PSCs special code has been added to the routines in **ukca_hetero_mod.F90**, and for other reactions there is code in **asad_hetero.F90**. Examples of this type are

```
rath_t('ClONO2    ','H2O       ','HOCl      ','HONO2     ','          ', &
'          ', 0.000, 0.000, 0.000, 0.000), &
...
rath_t('SO2       ','H2O2      ','NULL0     ','          ','          ', & !HSO3+H2O2(aq)
'          ', 0.000, 0.000, 0.000, 0.000),                                 &
```

To add new heterogeneous reactions you will need to append equivalent lines for the new reactions to the end of the **ratt_defs_scheme** array (increasing the array sizes accordingly), before adding code to either **ukca_hetero_mod.F90** or **asad_hetero.F90**.

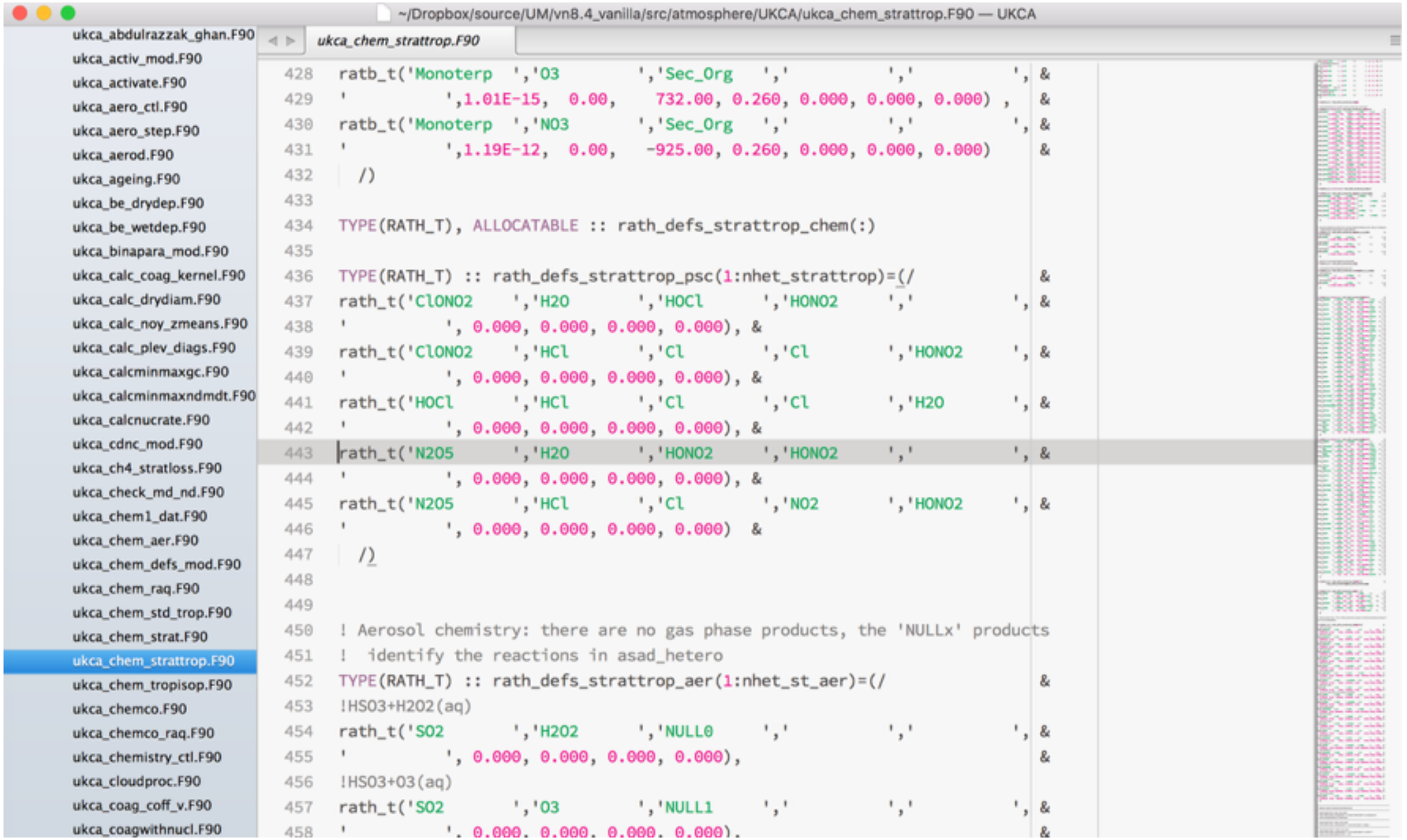
Increase the size of JPHK (and JPNR)

As with the bimolecular and termolecular reactions, you will also need to increase the values of two parameters that UKCA needs. These are

- **JPHK** is the number of heterogeneous reactions
- **JPNR** is the total number of reactions

These are set automatically in the UMUI (depending on what scheme is chosen), and are placed in the `&RUN_UKCA` namelist in **CNTLATM**. You will need to make a hand-edit to change these accordingly. The current values can be found by saving and processing the job, and then viewing the *CNTLATM* file in your `$HOME/umui_jobs/jobid` directory.

UKCA/ukca_chem_strattrop.F90



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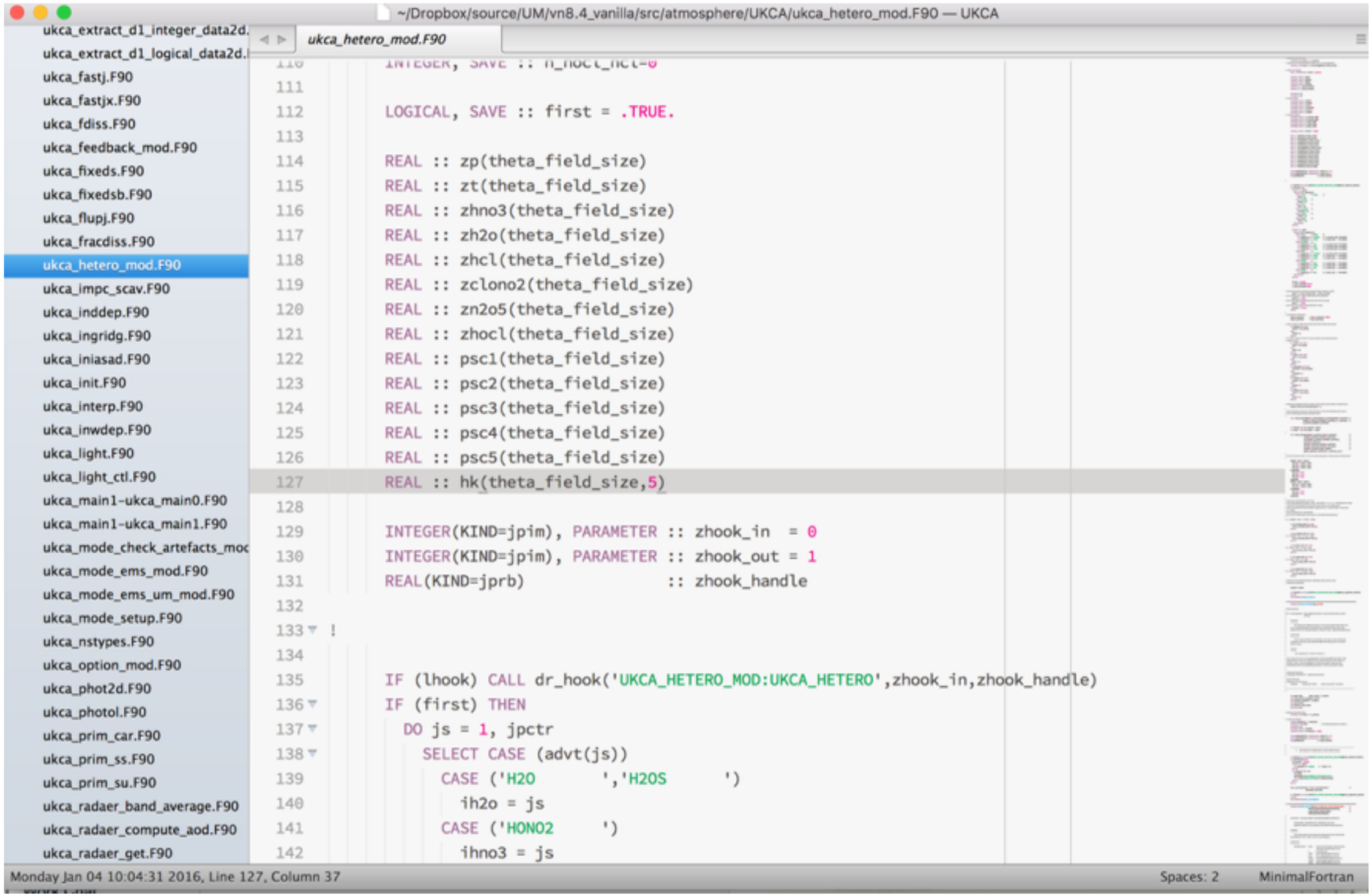

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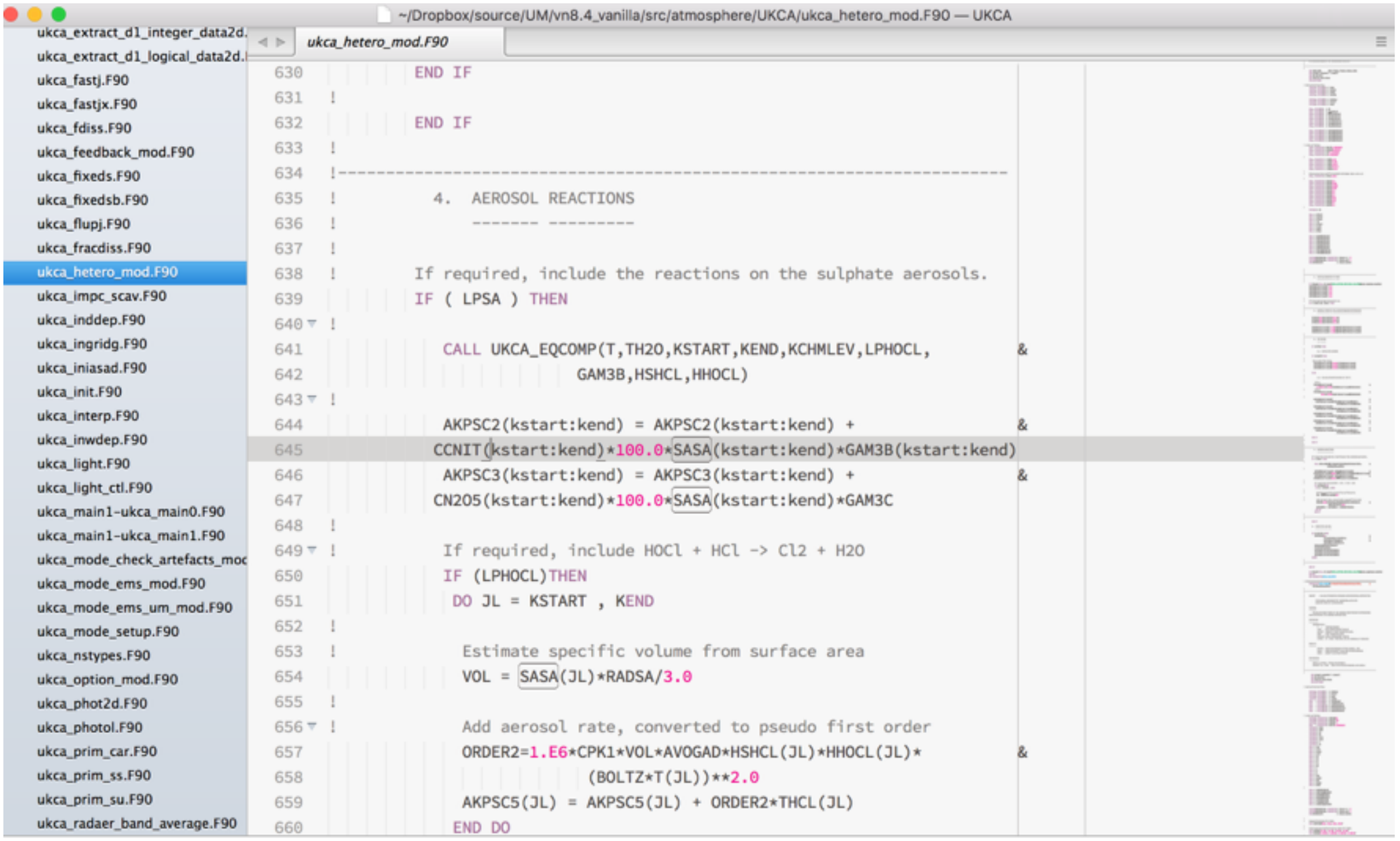
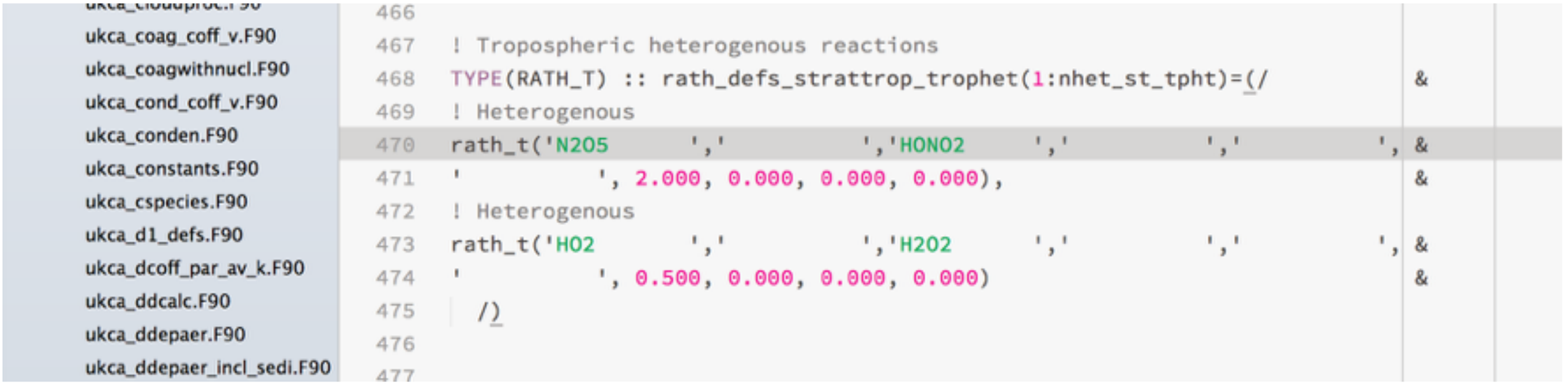


UKCA/ukca_chem_hetero_mod.F90

The MODE aerosol chemistry - slightly different

UKCA/ukca_chem_strattrop.F90



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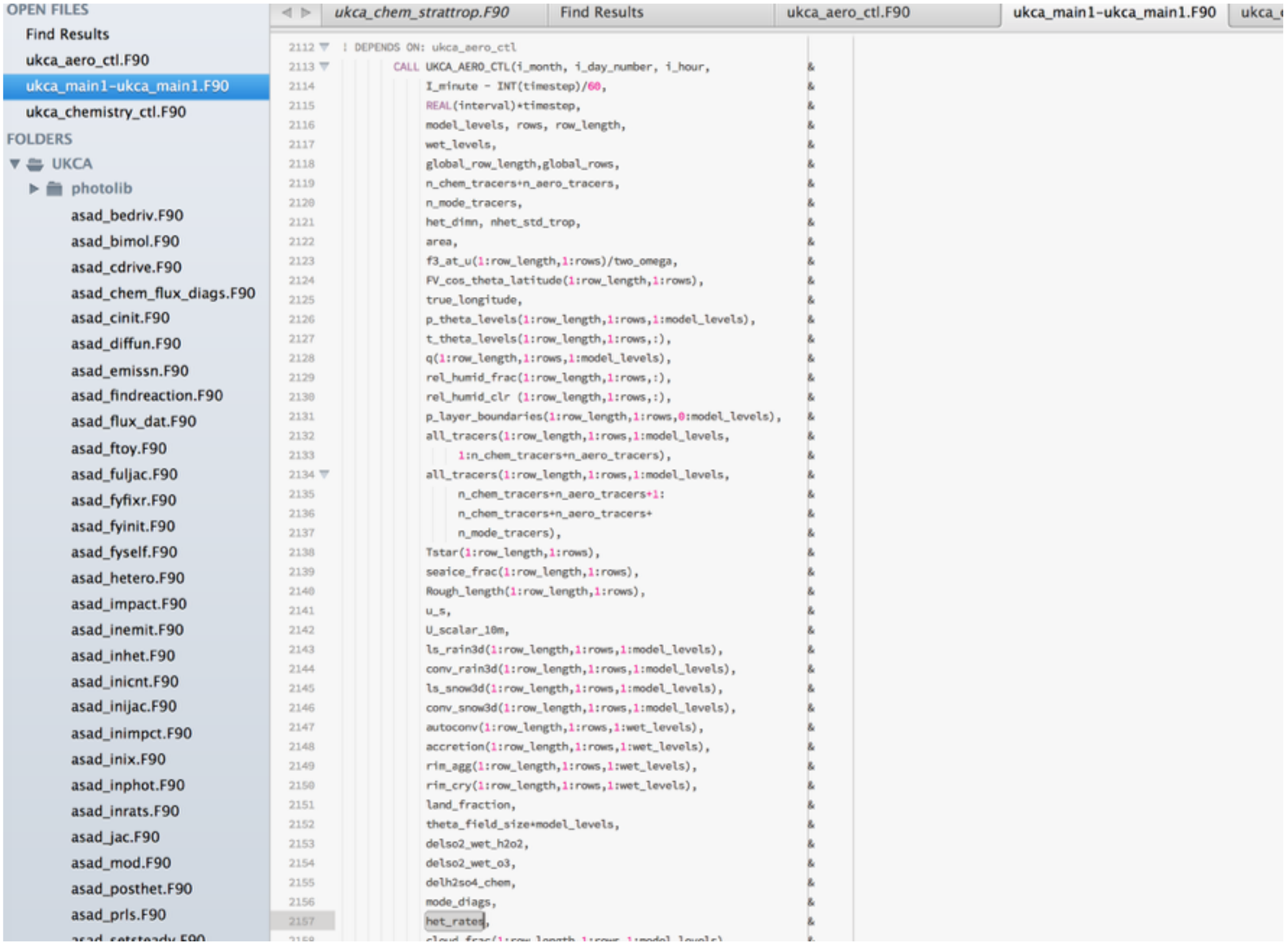

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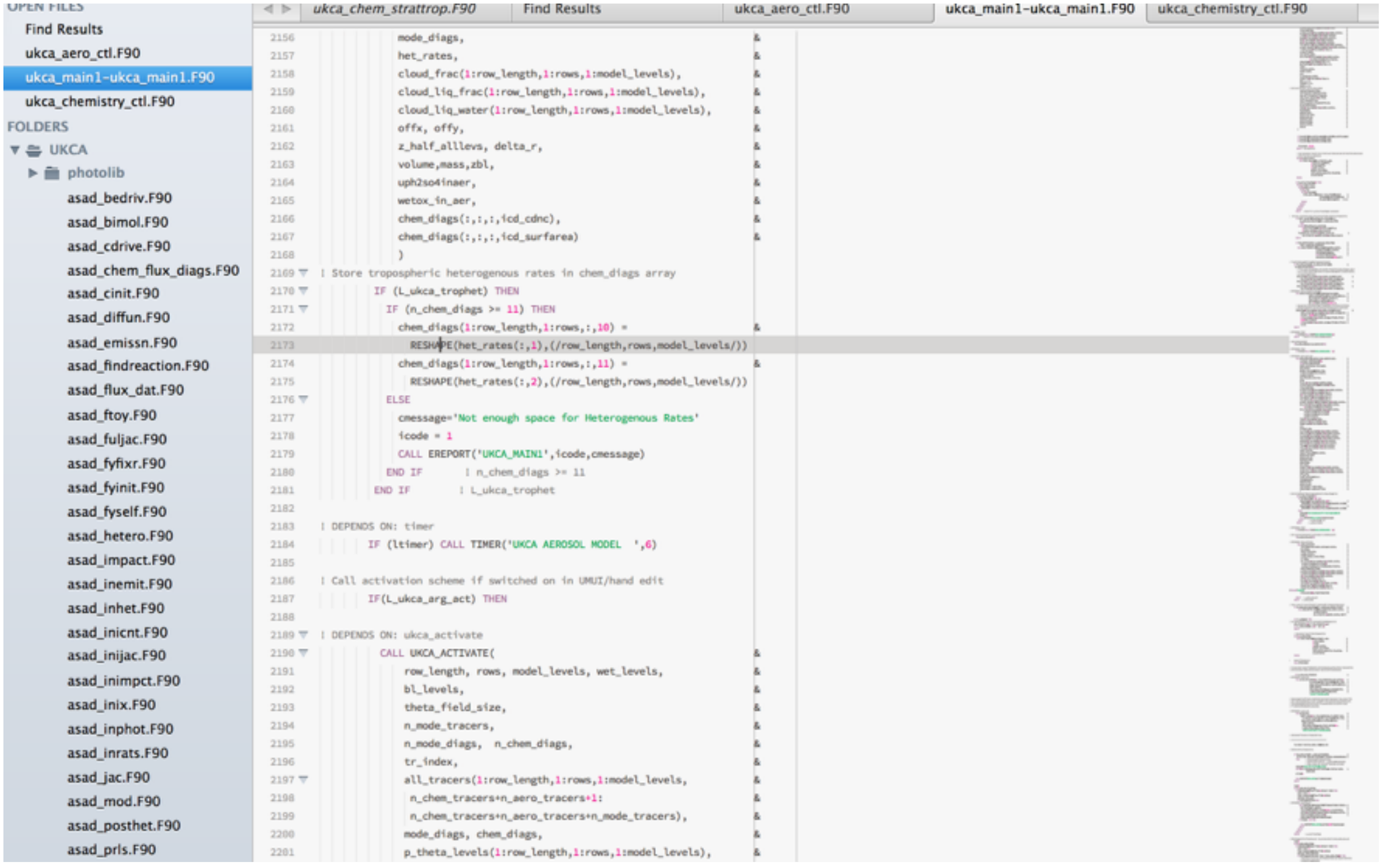

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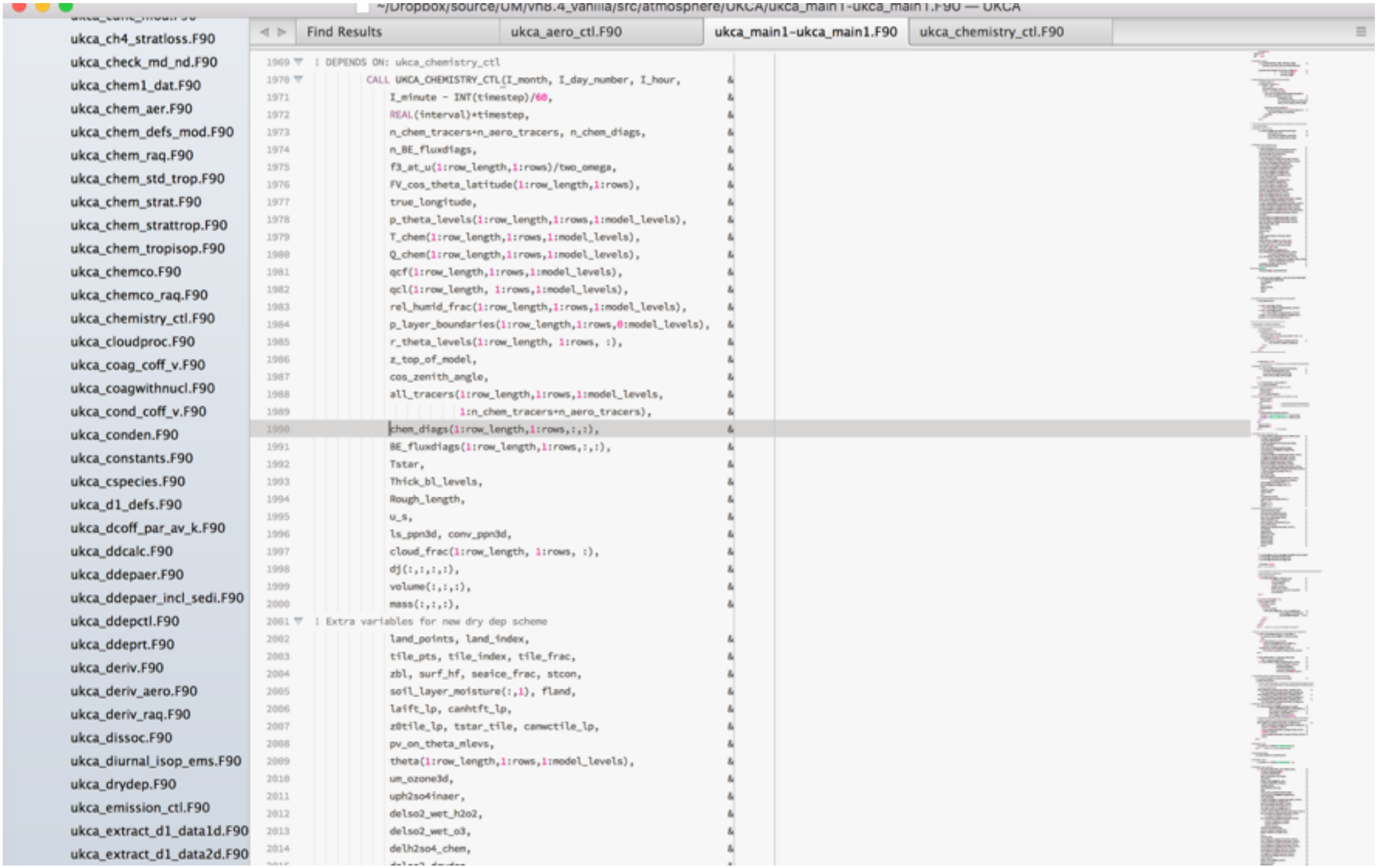

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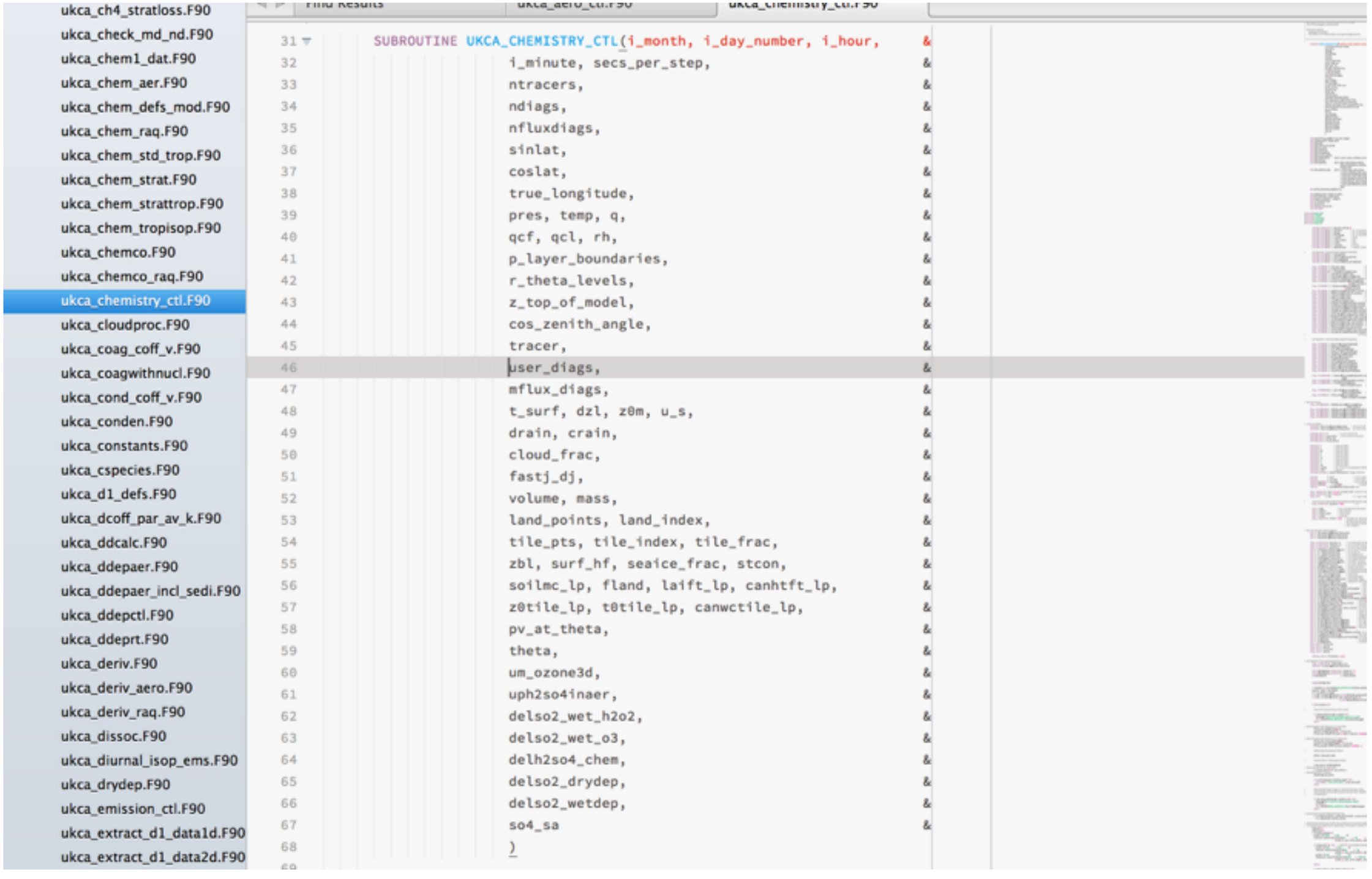

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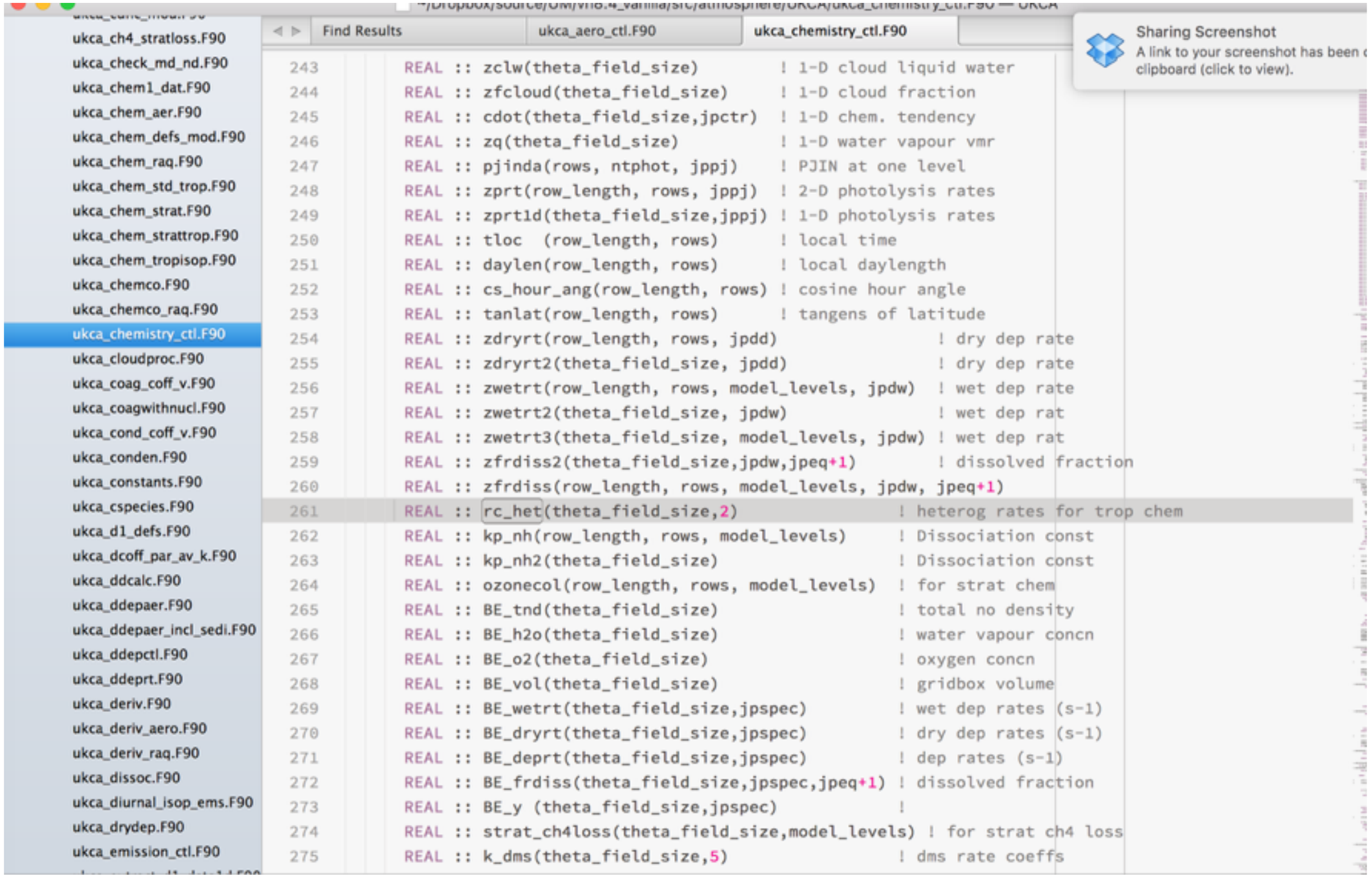

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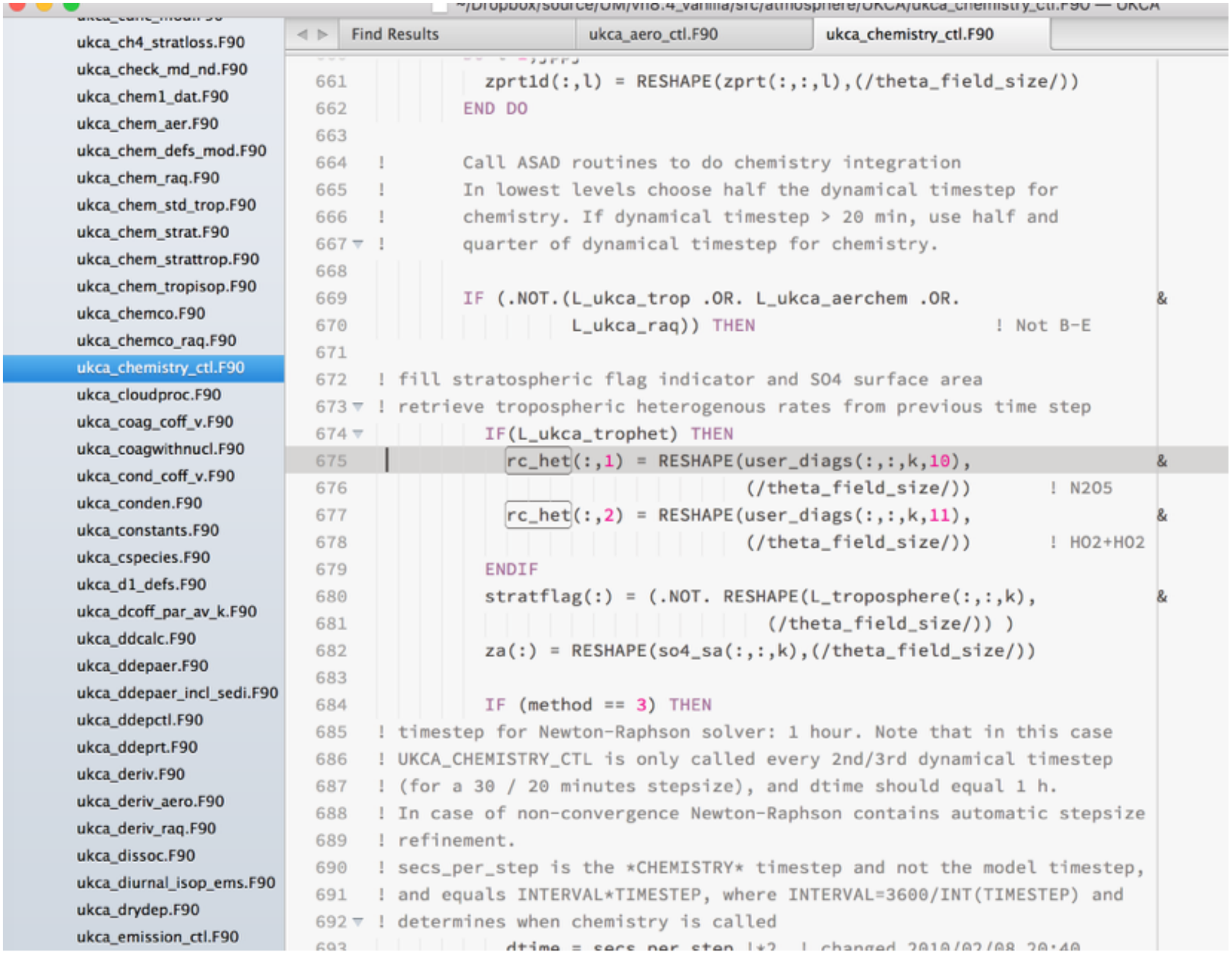

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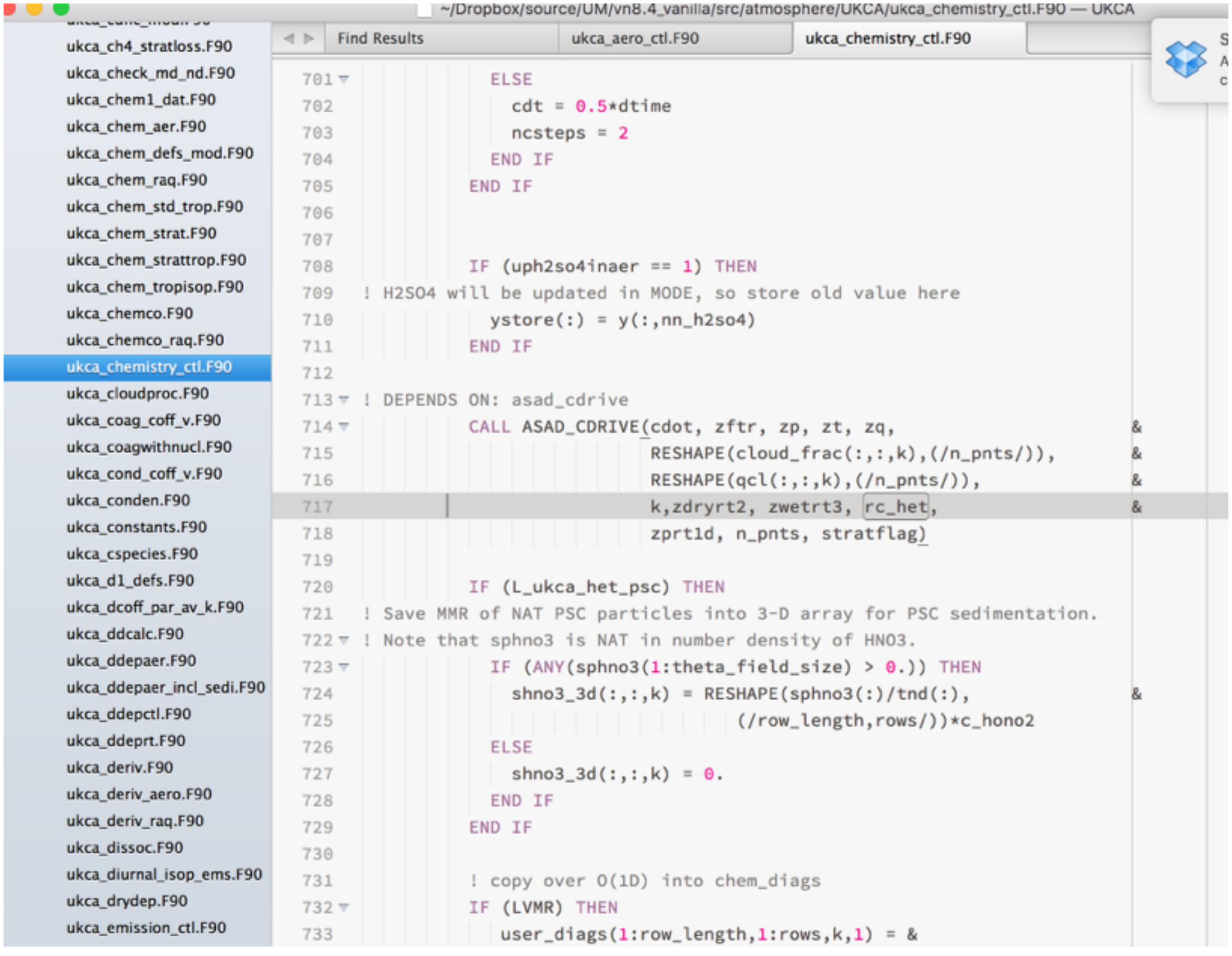

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